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theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...
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the following questions in the space
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equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value
of the coefficient x in this equation? b.
What is the molar mass of C_3H_4 ? c.

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What is the mole ratio of O_2 to H_2O in the above equation? d.

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equation: $C_3H_4(g) + x \cdot O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What is meant by ideal conditions

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following equation: $C_3H_4(g) + xO_2(g)$
 $? 3CO_2(g) + 2H_2O(g)$ 4 a. What is the
value of the coefficient x in this
equation? 40.07 g/mol b.

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Chapter 9 "I CANs"... ...represent and/or visualize chemical rxns from a math, micro and macro point of view ...use stoichiometry to convert moles &/or grams of one reactant &/or product into moles &/or grams of different reactants

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relationships between reactants and
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Reaction stoichiometry is the subject of
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N_2 combine with H_2 , what must be true about the quantity of H_2 for N_2 to be the limiting reactant? 4. If a reactionu0027s ...

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1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the

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Stoichiometry b. Theoretically, how many moles of NH_3 will be produced?
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to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? 40.07 g/mol b.

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reaction stoichiometry calculations start with a balanced chemical equation. This equation gives the

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in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g)$

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